

If you can't do this page you need to see me!
 Bonding Properties of Covalent Compounds

1. How do you recognize a covalent compound when you see one?

NON Metal - NON-metal

2. Ice is melted on a stove, what bond is breaking?

H-Bonding (Not Covalent)

3. What are the factors that create and amplify London Forces?

- Long chainy }
 - lots of e- } more polarizable

4. What are the factors that cause a substance to have a Dipole moment?

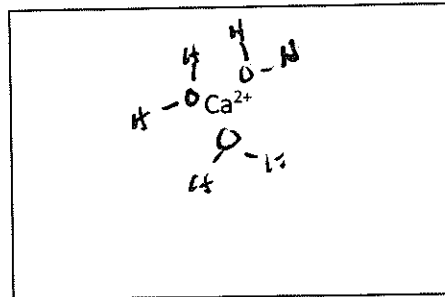
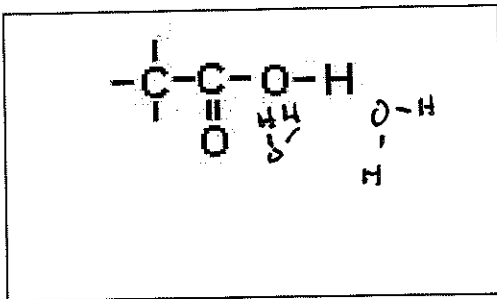
- difference in electronegativity
 - asymmetrical

5. Give two ways a substance can become polar?

- Transfer an e- to another atom
 - Have more e- on a side of a molecule

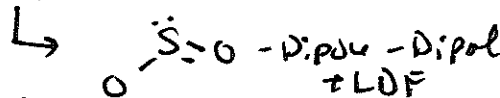
b)
 - Identical atoms Bonded
 - Symmetrical

6. In the box below draw a water molecule showing the dissolving process.



7. Arrange the following into increasing melting points/boiling points

a. ~~NO, O2~~, SO₂, I₂ — LDF — ↑ amounts



b. CH₄, NH₃, ~~CO₂, H₂O~~

↳ Condenses to liquid due to Hydrogen Bond
 ↳ only LDF